

Continue

Name _____ Date _____ Period _____

Periodic Table Worksheet

- Use a Periodic table to find the information asked for below:
1. What is the atomic number of:
 - Chlorine _____
 - Iron _____
 - Gold _____
 - Dtanium _____
 2. What is the Atomic mass of:
 - Chlorine _____
 - Iron _____
 - Dtanium _____
 3. How many protons do the following have?
 - Calcium _____
 - Copper _____
 - Iron _____
 4. How many electrons do the following have?
 - Gold _____
 - Copper _____
 - Iron _____
 5. Does mercury have more protons and electrons than tin?
 6. Is mercury a heavier element than tin?
 7. Does potassium have more electrons than iron?
 8. Does hydrogen have more electrons than Uranium?
 9. Which has more protons, sulfur or iodine?
 10. Which has more protons, iodine or silver?
 11. In the boxes below make Bohr models for each of the elements.
 - a. Determine how many electrons, protons, and neutrons there are in each atom.
 - b. Draw a Bohr model of each element using the number of electrons, protons, and neutrons
 - c. NOTE: The first energy level can only hold up to 2 electrons. The second energy level can hold up to 8 electrons.

_____ Electrons	_____ Electrons
_____ Protons HELIUM	_____ Protons OXYGEN
_____ Neutrons	_____ Neutrons

Chapter 3 & 6.4 Atomic/Ionic Radius Worksheet
Chemistry

Name: _____
Date: _____ Hour: _____

Atomic Radius (For help, see pp. 187 to 189 in the textbook.)

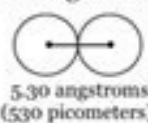
1. What is meant by the term atomic radius? _____
2. What is the trend across a row for atomic radius? _____
3. What is the trend down a column for atomic radius? _____
4. Predict which atom has a smaller atomic radius.
 - (a) magnesium or aluminum (circle one)
_____ same column or _____ same row (check one)
 - (b) strontium or barium (circle one)
_____ same column or _____ same row (check one)
 - (c) vanadium or niobium (circle one)
_____ same column or _____ same row (check one)
5. Predict which atom has a larger atomic radius.
 - (a) potassium or calcium (circle one)
_____ same column or _____ same row (check one)
 - (b) neon or argon (circle one)
_____ same column or _____ same row (check one)
 - (c) cesium or polonium (circle one)
_____ same column or _____ same row (check one)
6. Put the following elements in order from largest atom to smallest atom: carbon, fluorine, beryllium, and lithium _____
7. Put the following elements in order from smallest to largest: aluminum, iodine, boron _____

Atomic Radius Lab

Name _____ Hr _____

Background: The nucleus of an atom occupies a very small fraction of the atomic volume. The rest of the space is filled by the electrons traveling in the orbitals or electron "clouds". Just as it is difficult to determine the edge of a cloud, it is difficult to know when an atom "ends". The outer limit is fuzzy and variable.

Chemists have developed a method to determine the size of atomic radius. Using techniques such as X-ray diffraction and electron microscopy, chemists measure the distance between the nuclei of two adjoining atoms. Then they divide the distance in half and call it the radius of the atom. Consider that the location of the nucleus of an atom is stable. However, the edges of orbitals are "fuzzy" and can change according to conditions.



Determine the radius of the atom on the left.

Purpose: Determine the trend in atomic radius of atoms based on their locations in the periodic table.

Hypothesis:
a) How will the size of the atoms change moving down a column/group of the periodic table? Why?

b) How will the size of the atoms change moving across a period of the periodic table? Why?

Procedure:

Using the clay provided by your teacher, create a sphere to represent an atom of the periodic table. You will be assigned an element that you will produce to appropriate dimensions. Atomic radii are often given in angstroms (10^{-10} meters). You will make your sphere in centimeters. Remember, the diameter of the sphere will be double the radius!

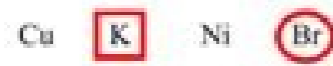
Atomic # & Symbol	Atomic Radius (Å)	Atomic # & Symbol	Atomic Radius (Å)	Atomic # & Symbol	Atomic Radius (Å)
1 H	0.37	12 Mg	1.60	33 As	1.20
2 He	0.31	13 Al	1.43	34 Se	1.19
3 Li	1.52	14 Si	1.18	35 Br	1.14
4 Be	1.12	15 P	1.10	36 Kr	1.12
5 B	0.85	16 S	1.03	37 Rb	2.48
6 C	0.77	17 Cl	1.00	38 Sr	2.15
7 N	0.75	18 Ar	0.98	52 Te	1.42
8 O	0.73	19 K	2.27	53 I	1.33
9 F	0.72	20 Ca	1.97	54 Xe	1.31
10 Ne	0.71	31 Ga	1.35	87 Fr	2.70
11 Na	1.86	32 Ge	1.22	88 Ra	2.20

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:



a. Explain why you made these choices: All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:



a. Explain why you made these choices: All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:



a. Explain why you made these choices: All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.

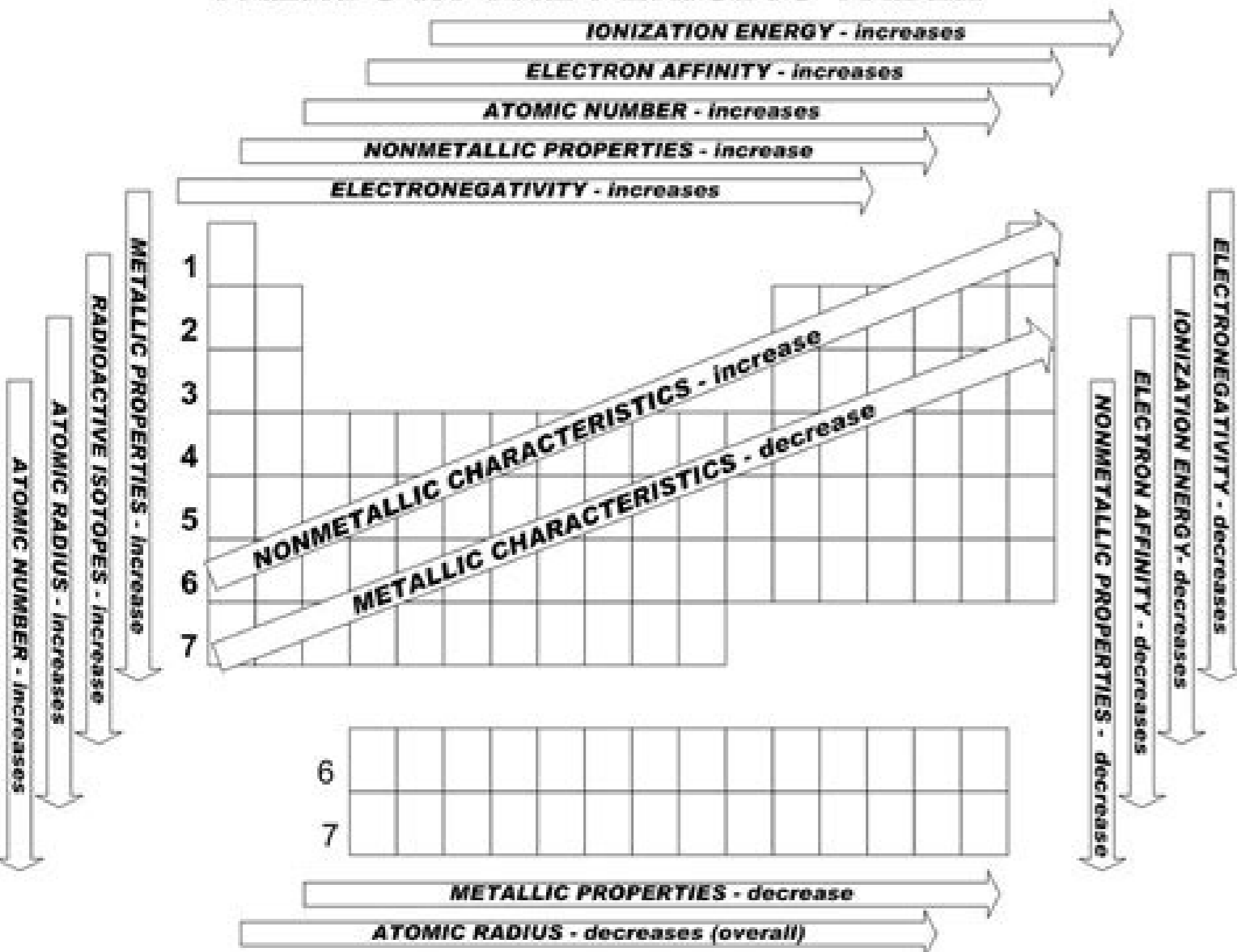
4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5.
 - a. O C Be Ne **Same Period**
 - b. Na Rb Fr H **Same Group**
 - c. Pb C Sn Si **Same Group**
 - d. Au W S Fr Ne Zn **Challenge**

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

5.
 - a. O C Be Ne **Same Period**
 - b. Na Rb Fr H **Same Group**
 - c. Pb C Sn Si **Same Group**
 - d. Au W S Fr Ne Zn **Challenge**

TRENDS IN THE PERIODIC TABLE



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