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4. Use a periodic table to fill in the missing information in the following table.

Name	Symbol	Atomic Number Z	Mass Number A	Number of Neutrons	Number of Electrons
oxygen	$^{16}_8\text{O}$	8	16	8	8
Manganese	$^{55}_{25}\text{Mn}$	25	55	30	25
Sulfur	$^{34}_{16}\text{S}$	16	34	18	16
Hydrogen	^1_1H	1	1	0	1
Hydrogen	^3_1H	1	3	2	1
Zinc	$^{64}_{30}\text{Zn}$	30	64	34	30
Zinc	$^{65}_{30}\text{Zn}$	30	65	35	30
Uranium	$^{238}_{92}\text{U}$	92	238	146	92
Krypton	$^{84}_{36}\text{Kr}$	36	84	48	36

Problems

1. The radius of a Cl nucleus is 4.0 fm, and the radius of a Cl atom is 100 pm. (1 fm = 1×10^{-15} m; 1 pm = 1×10^{-12} m). How many times larger is the diameter of the Chlorine atom than the diameter of the Chlorine nucleus?

1,000 times bigger

2. Identify two objects that have this same ratio of lengths.

1 ml of water and 1 L of water

3. How many times larger is the volume of the atom than the volume of the nucleus?

1,000 times bigger

Name: _____ Date: _____
 Class/Period: _____

Ion Practice

Part I - Directions: Read the questions below. Answer in the space provided.

1. An ion is a _____ particle formed from the loss/gain of a(n) _____.
2. A cation is a _____ charged particle formed from the _____ of an electron.
3. An anion is a _____ charged particle formed from the _____ of an electron.

Part II - Directions: Fill in the chart below using the information provided.

Ion Chemical Formula	Type of Ion	Type/# of particles lost/gained	# p ⁺	# e ⁻	# n ⁰	Atomic #	Mass #
4)		gain 1 electron			41		76
5) Cd ²⁺					68		
6)			27	24			60
7)				10		7	15

Part III - Directions: Fill in the chart below using the information provided.

Isotopic Notation	Isotope Name	Atomic #	Mass #	Oxidation #	# protons	# neutrons	# electrons
8)				-2		33	18
9) $^{60}_{28}\text{Ni}^{+1}$							
10)			58	+2	26		
11)		15	32	-3			

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